

Abstract

A comprehensive kinetic study of the title reaction was carried out at 65-80° with different concns. of HCl as catalyst. The solvent mixt. used was 1:1 (vol./vol.) methanol-water. The reaction obeyed an overall second-order rate law. The overall rate const. was resolved into rate consts. for the 2 steps. Activation energies and entropies were obtained. The exptl. and calcd. values of k at different temps. and concns. of acid catalyst agreed.